

Chemistry 111 Chapter 5 Cheat Sheet by Xyloswagg96 via cheatography.com/30111/cs/16152/

(V1/T1)=(V2/T2)

(P1(V1))/T1 = (P2(V2)/T2)

Converting ATM	
1 ATM=	1 atm=760 mmHg
Convert Unit of Hg to	mm
Multiply Unit	Unit X 1 atm/760 mm Hg
Divide Unit with 760 mm Hg	Unit/760

Convert kPa to atm	
1 ATM	1 ATM=101.3 kPa
Multiply	Unit X 1atm/101.3
Divide	Unit/101.3

Boyle's Law	
Boyle's Law	P1V1=P2V2

ımHg
760

Charles Law=

V=

Combined Gas Law

=	
Avogadro's Law	
Avogadro's Law is	(V1/n1)=(V2/n2)
Volume proportional	to number of moles
n=	quantity of gas in moles

Dalton's Law		
P_total=	P_a+P_b+P_c	

volume

Ideal Gas Law	
Ideal Gas Law=	PV=nRT
R=	0.08206
n=	Grams to Moles
T=	237
P=	usually 1



By Xyloswagg96 cheatography.com/xyloswagg96/

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