

Converting ATM

1 ATM= 1 atm=760 mmHg

Convert Unit of Hg to mm

Multiply Unit Unit X 1 atm/760 mm Hg

Divide Unit with 760 mm Hg Unit/760

Convert kPa to atm

1 ATM 1 ATM=101.3 kPa

Multiply Unit X 1atm/101.3

Divide Unit/101.3

Boyle's Law

Boyle's Law $P_1V_1=P_2V_2$

Charles Law

Charles Law= $(V_1/T_1)=(V_2/T_2)$

Combined Gas Law

Combined Gas Law $(P_1(V_1))/T_1=(P_2(V_2))/T_2$
=

Avogadro's Law

Avogadro's Law is $(V_1/n_1)=(V_2/n_2)$

Volume to number of moles proportional

n= quantity of gas in moles

V= volume

Dalton's Law

P_total= $P_a+P_b+P_c$

Ideal Gas Law

Ideal Gas Law= $PV=nRT$

R= 0.08206

n= Grams to Moles

T= 237

P= usually 1

C

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