

## Le-Chatelier Principle Cheat Sheet

by Tazz (Tonoya) via cheatography.com/162324/cs/44936/

Effect of concentration

Reactant--->product

Concentration increase

Product concentration increase

## Le-Chatelier Principle

At the equilibrium of reversible reaction, if any of the factors (temperature, pressure and concentration) is changed, the position of the equilibrium will shift in such a direction that the effect of change of factors is relieved.

Effect of temperature (exothermic reaction)				
A+B>C+D+Energy				
Or, A+B>C+D; ∇H=-ve				
Temperature increase	Temperature decrese			
Rate of forward reaction <rate backward="" of="" reaction<="" td=""><td>Rate of forward reaction&gt;Rate of backward reaction</td></rate>	Rate of forward reaction>Rate of backward reaction			
Equilibrium shift left ward	Equilibrium shift right ward			
Product concentration decrease	Product concentration increase			

Effect of pressure	
N <sub>2</sub> +3H <sub>2</sub> >2NH <sub>3</sub>	
Total mol more on left side	
Pressure increase	Pressure decrease
Rate of forward reaction>Rate of backward reaction	Rate of forward reaction <rate backward="" of="" reaction<="" td=""></rate>
Equilibrium shift right ward	Equilibrium shift left ward
Product concentration increase	Product concentration decrease

Pressure decrese
Rate of forward reaction>Rate of backward reaction
Equilibrium shift right ward
Product concentration increase

Effect of pressure	
H <sub>2</sub> +l <sub>2</sub> >2HI	
Total mol equal on both side	
Pressure increase	Pressure decrease
No effect	No effect

## Effect of pressure

Only applicable for gasses as pressure doesn't affect solid or liquid.



By **Tazz** (Tonoya) cheatography.com/tonoya/

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Rate of forward reaction>Rate of backward reaction	Rate of forward reaction <rate backward="" of="" reaction<="" th=""></rate>			
Equilibrium shift right ward	Equilibrium shift left ward			
Product concentration increase	Product concentration decrease			
Effect of temperature (endothermic reaction)				
A+B>C+D-Energy				
A+B>C+D-Energy Or, A+B>C+D; ∇H=+ve				
0,7	Temperature decrese			
Or, A+B>C+D; ∇H=+ve	Temperature <i>decrese</i> Rate of forward reaction <rate of<="" td=""></rate>			
Or, A+B>C+D; ∇H=+ve Temperature <i>increase</i>				

Concentration decrease

Product concentration decrease