Seperation Techniques

S+L - Soluble	Crystallization
S+L Insoluble	Filtration, Decantation
L+L Miscible	Distillation
L+L Immiscible	Separating funnel
G+G	Fractional distillation

Structure 1

Elements	An element is a pure substance made up of only one type of atom
Molecule	Molecules of an element contains the same kind of atoms
Compound	A pure substance made up of two or more types of atoms that are chemically bonded together in a fixed ratio.
Mixture	A combination of two or more elements or compounds that are not chemically combined.

Mixtures

Homogeneous	It has a uniform or even distribution of constituent particles.
Heterogeneous	It has a non-uniform or uneven distribution of constituent particles.
Colloid	a mixture of two substances, where one substance is made up of microscopically dispersed particles that are suspended in the other substance
Sols	A colloidal system in which solid particles are dispersed in a liquid
Gels	when a sol as a rigid form, it is referred to as a gel.
Emulsions	a colloidal dispersion of one liquid in another when both

Celsius to Kelvin

0°C + 273.15 = 273.15K

Kinetic theory of matter

KTP States that all matter is composed of particles which have intermolecular space between them, attract each other witch force and are in continuous random motion

1.2.1

Particle	Relative Mass	Relative Charge
Proton	1 amu	+1
Neutron	1 amu	0
Electron	1/1836 amu	-1

Radioisotopes

Isotopes differ in the number of neutrons resulting in different mass numbers

Radioactive isotopes are used in nuclear medicine for:

- Diagnostics
- Treatment
- Research
- Tracers
- Geological clocks

By **shachi**
cheatography.com/shachi/

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