

Chem Rules and Formulas Cheat Sheet by Schutte via cheatography.com/30544/cs/9121/

Chapter 1 Formulas		Chapter 4 Oxidation Numbers				Chapter 4 Solubility Rules (cont)			
$T(^{\circ}C) = T(K) - 273.15$ $T(^{\circ}F) = 9/5 [T(^{\circ}C)] + 32$		element by	0	0			CH3CO2-	PO4 ³⁻	Na+, K+,
Density = mass/volume		Group 1A	+1	+1					NH4+
average atomic weight = (exact mass of		Group 2A	+2				CIO3^-		
isotope) (fractional abundance of isotope) +		Halogen	usually	usually -1 / positive with Oxygen			CIO4-		
		Monatomic I	lon ion charge			Pb ²⁺ ,	CI-		
Chapter 1 Conversions		Н	+1 with	+1 with nonmetals / -1 with metals					
		0	usually	-2 / -1 in pero	xide (O2)	Hg2 ²⁺			
1 in	2.54 cm	F	ALWAY	'S -1		Pb ²⁺ ,	Br-		
1 mi	5280 ft					Ag+, Hg2 ² +			
Diatomic Molecules		Chapter 4 Solubility Rules				Pb ²⁺ ,			
H2	olecoules	exceptions	soluble ionic	insoluble ionic	exceptions	Ag+,	l-		
N2			compounds	compouns		Hg2 ²⁺			
O2			Na+	OH-	Na+, K+,	Ca+, Sr+,	F-		
F2				Ba ² + O ² - Na+, K+,		Ba ²⁺ ,			
CI2			K+						
Br2				_	Ba ²⁺	Ca+, Sr+, Ba ²⁺ ,	SO4^2-		
12			NH4+	S ² -	Na+, K+, NH4+, Ba+				
Polyatomic lons			NO3-	CO3 ² -	Na+, K+, NH4+	Pb+			
D SOUTH					INIT				



By **Schutte** cheatography.com/schutte/

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