Alkenes

Naming: ends in *-ene*
e.g. but-2-ene

Physical Properties

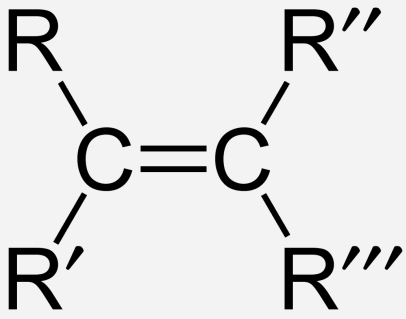
~ Insoluble in water
~ Melting point increases with Carbon chain length
~ Low melting point

Preparation

Can be formed using:
Alkane -> Alkene (H2/Pt)
Alcohol -> Alkene (H+/H2O)
Haloalkane -> Alkene (HX)
Alcohol (diol) -> Alkene (H+/MnO4)
Haloalkane -> Alkene (X2)

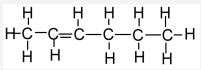
X = halogen

Functional Group



-C=C-

Hex-2-ene



Isomers

trans	cis
Methyl groups on opposite sides of double bond	Methyl groups on same side of double bond

see diagram below

Trans isomer

Cis isomer

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