

Redox Reaction

oxidation - loss of electrons reduction - gain of electrons

oxidation number increases oxidation number decreases

the thing that is oxidized is the thing that is reduced is
the reducing agent the oxidizing agent

Oxidation Number Rules

group 1: always +1

group 2: always +2

aluminum: +3

fluorine: always -1

halogens: -1

hydrogen w/a: nonmetal: +1 | metal: -1

oxygen: -2 | peroxides -1 | superoxides -1/2

halogens: -1 except with oxygen and transitional metals

nonmetals: usually negative oxidation states

metals: usually positive oxidation states



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Page 1 of 1.

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