

Orbitals + Hybridization

S orbital	Sphere, low energy, one orbital, holds up to 2e
P orbital	Dumbell, high energy, 3 orbitals, can hold up to 6e
Hybridization	The hybridization is $sp(\text{# different e- fields} - 1)$

Lewis structure + Formal charge

Molecular geometry

Sigma bonds	all single bonds, stronger, usually allows rotation, axial overlap
Pi bonds	double and triple bonds, slightly longer, restricted rotation creating cis and trans



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Not published yet.

Last updated 30th November, 2025.

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