

Orbitals + Hybridization

S orbital	Sphere, low energy, one orbital, holds up to 2e
P orbital	Dumbell, high energy, 3 orbitals, can hold up to 6e
Hybridization	The hybridization is $sp(\# \text{ different e}^- \text{ fields} - 1)$

Lewis structure + Formal charge

Molecular geometry

Sigma bonds	all single bonds, stronger, usually allows rotation, axial overlap
Pi bonds	double and triple bonds, slightly longer, restricted rotation creating cis and trans



By NoelleEvelyn

cheatography.com/noelleevelyn/

Not published yet.

Last updated 30th November, 2025.

Page 1 of 1.

Sponsored by **ApolloPad.com**

Everyone has a novel in them. Finish Yours!

<https://apollopad.com>