Chemistry Final Cheat Sheet

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Sig Fig Rules

To determine the number of significant figures in a number use the following 3 rules:

1. Non-zero digits are always significant

2. Any zeros between two significant digits are significant

Example: .500 or .632000 the zeros are significant .006 or .000968 the zeros are NOT significant

For addition and subtraction use the following rules:

1.Count the number of significant figures in the decimal portion ONLY of each number in the problem

2.Add or subtract in the normal fashion

3.Your final answer may have no more significant figures to the right of the decimal than the LEAST number of significant figures in any number in the problem.

For multiplication and division use the following rule:

1.The LEAST number of significant figures in any number of the problem determines the number of significant figures in the answer. (You are now looking at the entire number, not just the decimal portion) This means you have to be able to recognize significant figures in order to use this rule

Example: 5.26 has 3 significant figures -> 6.1 has 2 significant figures

Method for conversions

#unit1 x #unit (converting to) / #unit1	tera	10^12
cancel like units	giga	10^9
then multiply and divide then you get your answer with new units	mega	10^6
	kilo	10^
	centi	
	milli	
	micro	
	nano	
	pico	

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