

Conceptual Questions

What is the Big Idea of Chapter 13?

Molecules can be drawn using simplified models. These models can be modified in order to account for bond angles and shape.

Explain VSEPR

VSEPR is the idea that molecular shape is caused by electrons minimizing the repulsion between them to create a more stable molecule.

What causes different molecular shapes?

Electron repulsion causes different shapes.
















When is a molecule polar?

A molecule is polar when there are some regions of the molecule that have a stronger pull on the electrons.

What is formal charge?

Formal charge is the "net worth" of an atom. It measures how much it has versus how much it spent in bonding.

MOLECULAR GEOMETRY

Steele No.	Basic Geometry: 0 lone pairs	VSEPR Geometries			
		1 lone pair	2 lone pairs	3 lone pairs	4 lone pairs
2	 Linear				
3	 Trigonal Planar	 Bent < 120°			
4	 Tetrahedral	 Trigonal Pyramidal Bent < 109.5°	 Bent < 109.5°		
5	 Trigonal Bipyramidal	 Seesaw or Ureteric	 T-shape	 Linear	
6	 Octahedral	 Square Pyramidal	 Square Planar	 T-shape	 Linear

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Not published yet.

Last updated 7th June, 2017.

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