Changing state

The temperature will remain constant during the change of state.
The melting and freezing point are the same temperature.

Mixtures

Homogeneous mixtures
Heterogeneous mixtures

Group 1

Alkali
Silvery-white coloured
Metals
Relatively low boiling points

They react with water to produce hydrogen gas and heat.
The heat produced can ignite (burn) the hydrogen gas produced.
As you move down the group, the reactions become more violent.

Group 2

Alkaline
Relatively high boiling points
The reactions are less violent than group 1 elements.

Group 17

Halogens
diatomic
very reactive with metals
F2 and Cl2 pale green gases
Br2 red brown liquid
I2 grey solid

Group 18

nobel gases
unreactive

Covalent Compounds

non-metal atoms chemically bonded together

Ionic Compounds

monatomic ion	polyatomic ion
single atom that has a charge	group of atoms with a charge

Predicting bond type

Metals and non-metals ionic bonds
Non-metals covalent bonds

Reactions of Metals

Reactions of Metals	Reactions of Acids and Bases
Metal + Oxygen Metal oxide Metal + Water Metal hydroxide + Hydrogen Metal + Acid Metal salt + Hydrogen Gas	Acid + Carbonate Metal salt + Water + Carbon dioxide Acid + Base Metal salt + Water

Metal + Oxygen Metal oxide
Metal + Water Metal hydroxide + Hydrogen
Metal + Acid Metal salt + Hydrogen Gas

Corrosion

name 3 methods to reduce corrosion

> Painting Tin coating Chrome plating Enamelling Plastic coating Galvanizing (Zinc coating) Anodizing Alloying

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