Acids

Acids+metals→Salt+Hydrogen gas
Metal Carbonate/Metal Hydrogencarbonate+Acid→Salt+CO2+Water
Acid+Base→Salt+Water
Metal Oxide+Acid→Salt+Water

Acid or base in a water solution

When an acid or a base is dissolved in water, they get dissociated into ions.
For example,
When hydrochloric acid is dissolved in water, it get dissociated into ions such as protons (H+ ions) and Cl- ions as follows:
$HCl + H_2O \rightarrow H_3O^+ + Cl^-$
As there is an increase in the protons in the aqueous solutions, the solution is acidic in nature.
Similarly, when NaOH is dissolved in water, it get dissociated as,
$NaOH + H_2O \rightarrow Na^+ + OH^- + H_2O$
As there is an increase in the hydroxyl ions in the solution, the solution is basic in nature.
Basic aqueous solution is called alkali.

How strong are acids and base solutions

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