

Pure Substances

a single substance or matter which cannot be separated into other kinds of matter by any physical changes

There are two types of pure substances *elements* and *compounds*

ELEMENTS = pure substance made up of one type of atoms only

COMPOUNDS = pure substance containing two or more types of elements which are combined together in a fixed proportion by mass

Elements

SOLID ELEMENTS = copper silver gold potassium carbon(diamond and graphite) iodine phosphorus etc

LIQUID ELEMENTS = only mercury and bromine exist at room temperature. Gallium and Cesium become liquids at a temperature of 302K and 303K. These are slightly higher than the room temperature 298K.

GASEOUS ELEMENTS = eleven elements exist in gaseous state at room temperature. These are hydrogen oxygen chlorine fluorine helium argon neon nitrogen xenon krypton and radon

Metals

Metals *mercury is liquid at room temperature.*
are *Gallium and Cesium*
solid *become liquid at*
at *room temperature slightly*
room *above the room*
temper *temperature*
ature

Metals (cont)

The atoms are very closely packed in space. This arrangement is known as crystal lattice. Lattice varies from metal to metal.

Metals have shiny surfaces. they generally have silver-grey or golden-yellow surfaces. This property is known as lustre.

Metals are good *Copper*
conductor of heat *and*
and electricity. *aluminium*

Metals are *Potassium*
generally quite *and*
hard *sodium*

Metals are malleable

Metals are ductile

Metals are sonorous

Metals generally have high melting and boiling points

Non-Metals

are either gases or *bromine*
solids at room
temperature

vary in colour. solids **crystals*
have dull surfaces *of iodine*

mostly poor *Graphite*
conductors of heat
and electricity

Non-Metals (cont)

Most of them are *diamond*
quite soft and have
less densities than
metals

non malleable and non ductile in nature

have very low melting and boiling points

Compounds

Inorganic *common*
compounds = have *salt*
been obtained from *marble*
non living sources *washing*
such as rocks and *soda*
minerals *baking*
soda
carbon
dioxide
ammonia
etc

Organic compounds *methane*
= obtained from *ethane*
living beings such *propane*
as plants and *alcohol*
animals. contain *sugar*
carbon as their *proteins*
essential *oils fats*
component. also *etc*
called carbon
compounds

Acids Sulphuric
acid,
hydroc-
holoric
acid and
nitric
acids

Compounds (cont)

Bases Sodium hydroxide,
Potassium hydroxide
and calcium hydroxide

Salts sodium chloride,
calcium nitrate and
zinc sulphate

Compounds Properties

composed of the same elements combined in a fixed ratio by mass to form molecules

a pure compound is homogeneous in nature

a chemical compound is formed as a result of chemical reaction between constituent elements

properties of a compound are different from the elements from which it is formed

constituents of a chemical compound cannot be separated mechanically

formation or decomposition of compounds involves energy changes

Physical change

brings change in the physical state of matter under suitable conditions

Properties changes the interparticle forces or the interparticle spaces



Physical change (cont)

no change in the composition of the substances

do not change their main characteristics

no new substance is formed

change is temporary and can be reversed by reversing the conditions which bring about the change

no energy change generally occurs during a physical change

Chemical change

brings a change in the chemical composition of the matter

generally there is an exchange in constituents and new substances are formed

Properties Physical state of the substance may or may not change

always a change in the physical composition of the substances undergoing chemical changes

Chemical change (cont)

change in the characteristics of the substance involved

new substance is always formed

chemical changes are permanent in nature and cannot be easily changed

energy reactions always occur in chemical reactions responsible for these changes

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Solutions (cont)

aerated drinks

air

copper sulphate in water

dilute hydrochloric acid

brass

bronze

solder

Properties of a solution

homogenous in nature

all components are present in the same phase

particles cannot be seen by naked eye or ordinary microscope

solution particles can pass through the fine pores of the filter paper

the components do not settle down if left undisturbed for a very long time. this shows that a solution is quite stable in nature

particles do not scatter a beam of light

a saturated solution becomes unsaturated upon heating

a solution in which water acts as the solvent it is known as aqueous solution while the solution in which the solvent is another liquid it is non aqueous solution

Solutions

a homogenous mixture of two or more non reacting substances

Types of solution:

Solid solution solid acts as the solvent

Liquid solution liquid acts as the solvent

Gaseous solution gas acts as the solvent

Only a mixture of miscible liquids is a solution. In case, they do not mix with each other and form separate layers they are known as emulsion

in the homogenous mixture, the particle size is 1nm in diameter

alloys are homogenous mixture of two or more metals or non metals

Examples of solution sugar in water

iodine in alcohol



Saturated solution

a solution becomes saturated if the solute starts separating at the bottom of the container in which the solution is being prepared at a given temperature

a saturated solution becomes super saturated upon cooling

Suspensions and Colloidal Solutions

a heterogenous mixture in which the solid particles are spread throughout the liquid without dissolving in it. they settle as precipitate if the suspension is left undisturbed for a while

Properties heterogenous nature

particle size is more than 100nm

particles can be seen with naked eyes

solid particles can be easily separated through ordinary filter papers

particles are unstable as they settle down if the suspension is left undisturbed. this is known as precipitate

Colloidal Solutions

heterogenous in nature but have smaller size of particles which are undistributed .It ranges between 1nm to 100nm

Properties appear to be homogenous but are hetero- genous

are a two phase system

particles pass through ordinary filter papers

particles carry charge

particles follow a zig- zag path

the scattering of the beam of light by the dispersed phase of particles is known as Tyndall effect

colloidal solutions in which only liquids participate are known as emulsions

Mixtures

the combination of two or more substances which are physically mixed and are not chemically combined with each other and may be also present in any proportions

Homogenous Mixture the components are mixed uniformly mixed without any clear boundary of seperation

sodium chloride and sugar in water . Air is a homogenous mixture of gases

Hetero- genous Mixture the components do not have uniform composition and also have visible boundaries of separation between the constituents

sand and common salt. Oil and water