Electronegativity

Increases Across the Table

- The Atomic number increases
- The number of protons increases
- The valence electrons all go in the same shell
- The nuclear charge is effective
- The attraction between the protons and electrons increases

= Therefore the electronegat-ivity increases

Electronegativity

Decreases Down the Table

- The number of shells increases
- The electrons are further away from the nucleus
- The amount of shielding increases
- The attraction between the protons and electrons decreases

= Therefore the electronegatitvity decreases

Atomic Radii

Increases Down the Table

- Number of protons increase
- The positive charge of the nucleus increases
- The number of shells increase
- Shielding is increased
- The attraction between electrons and protons decrease

= Therefore the atomic radii increases

Atomic Radii

Increases Down the Table

- The Atomic number increases
- The number of protons increases
- The positive charge of the nucleus increases
- Valence electrons are added to the same shell
- The protons and electrons are attracted
- The valence shell is pulled towards the nucleus

= The atoms are smaller

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