Cheatography

Electronegativity

Trends of the Periodic Table Cheat Sheet by holdover247 via cheatography.com/20637/cs/3358/

Increases Across the Table	Increases Down the Table
- The Atomic number increases	- Number of protons increase
- The number of protons increases	- The positive charge of the
- The valence electrons all go in	nucleus increases
the same shell	- The number of shells increase
- The nuclear charge is effective	- Shielding is increased
- The attraction between the	- The attraction between electron
protons and electrons increases	and protons decrease
= Therefore the electronegativivity	= Therefore the atomic radii
increases	increases
Electronegativity	Atomic Radii
Decreases Down the Table	Increases Down the Table
- The number of shells increases	- The Atomic number increases

Atomic Radii

from the nucleus - The amount of shielding

increases

- The attraction between the protons and electrons decreases

- The electrons are further away

= Therefore the electronegatitvity decreases

- The number of protons increases
- The positive charge of the
- nucleus increases
- Valence electrons are added to the same shell
- The protons and electrons are attracted
- The valence shell is pulled towards the nucleus
- = The atoms are smaller



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