

### Electronegativity

#### Increases Across the Table

- The Atomic number increases
- The number of protons increases
- The valence electrons all go in the same shell
- The nuclear charge is effective
- The attraction between the protons and electrons increases
- = Therefore the electronegativity increases

### Atomic Radii

#### Increases Down the Table

- Number of protons increase
- The positive charge of the nucleus increases
- The number of shells increase
- Shielding is increased
- The attraction between electrons and protons decrease
- = Therefore the atomic radii increases

### Electronegativity

#### Decreases Down the Table

- The number of shells increases
- The electrons are further away from the nucleus
- The amount of shielding increases
- The attraction between the protons and electrons decreases
- = Therefore the electronegativity decreases

### Atomic Radii

#### Increases Down the Table

- The Atomic number increases
- The number of protons increases
- The positive charge of the nucleus increases
- Valence electrons are added to the same shell
- The protons and electrons are attracted
- The valence shell is pulled towards the nucleus
- = The atoms are smaller

