# Trends of the Periodic Table Cheat Sheet

by holdover247 via cheatography.com/20637/cs/3358/

## Electronegativity

**Increases Across the Table**

- The Atomic number increases
- The number of protons increases
- The valence electrons all go in the same shell
- The nuclear charge is effective
- The attraction between the protons and electrons increases

= Therefore the electronegat-ivity increases

## Electronegativity

**Decreases Down the Table**

- The number of shells increases
- The electrons are further away from the nucleus
- The amount of shielding increases
- The attraction between the protons and electrons decreases

= Therefore the electronegatitvity decreases

## Atomic Radii

**Increases Down the Table**

- Number of protons increase
- The positive charge of the nucleus increases
- The number of shells increase
- Shielding is increased
- The attraction between electrons and protons decrease

= Therefore the atomic radii increases

## Atomic Radii

**Increases Down the Table**

- The Atomic number increases
- The number of protons increases
- The positive charge of the nucleus increases
- Valence electrons are added to the same shell
- The protons and electrons are attracted
- The valence shell is pulled towards the nucleus

= The atoms are smaller

By **holdover247**

cheatography.com/holdover247/

Not published yet.
Last updated 12th May, 2016.


Sponsored by **Readable.com**
Measure your website readability!
https://readable.com