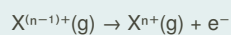


Ionisation energy



For the n^{th} ionisation energy

Number of particles

Number of moles \times Avogadro's constant
(6.022×10^{23})

Number of moles

Mass of substance (g) / Mr

Concentration (mol dm^{-3}) \times Volume (dm^3)

Ideal gas equation

$$pV = nRT$$

p=pressure (Pa)

V=volume (m^3)

n=number of moles

R= $8.31 \text{ J K}^{-1} \text{ mol}^{-1}$

T=temperature (K)

Arrhenius Equation

$$k = Ae^{-\frac{E_a}{RT}}$$

C

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