

### Finals

#### Acids and bases-Definitions

##### Acids and Bases

Arrhenius Acid: produces  $H^+$  ions in water

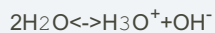
Arrhenius Base: Produced  $OH^-$  ions in water

Bronsted Acid: Proton ( $H^+$ ) donor

Bronsted Base: Proton ( $H^+$ ) acceptor/receiver

Strong Acids: Donate 100% of the  $H^+$  ion in water      Goes to completion

Weak Acids: Donates less than 100% on  $H^+$  ions in water



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By [danielle\\_ha\\_ha](https://cheatography.com/danielle-ha-ha/)  
[cheatography.com/danielle-ha-ha/](https://cheatography.com/danielle-ha-ha/)

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