

Energy guide Cheat Sheet by ChemIsForSheeit via cheatography.com/141751/cs/31083/

Constants

h= 6.63×10^{-34} **J*s** c= 3×10^{8} **m/s**

h is Planck's Constant c is speed of light

Purpose for flame test

To identify a metal by its color

When is color emitted?

When the electron comes back to its groundstate

Harmful and little effect radiations

Gamma Rays Long Radio waves

IR or Infrared

X rays Microwaves

C=VA (A is lambda)

UV or UltraViolet

E=hV

E is energy

V is frequency

A is lambda

How an electron emits a photon

Heat or electricity is applied to an atom, and the electron absorbs the energy and then gets excited; leaves its ground state, and comes back to its nucleus due to it being attracted, and then releases energy; Photon

Conversions

1m=1x10⁹nm 1kHz=1x10³Hz 1mHz=1x10⁶Hz

mole of a photon

(6.02x10²³)(6.63x10⁻³⁴)(V)
Multiply this exactly how this is once you get your **V** and you will get your mole of a photon for the problem.



By ChemIsForSheeit

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