

Acids and Bases

Strong Acids	Strong Bases
HClO ₄	LiOH
H ₂ SO ₄	NaOH
HI	KOH
HBr	Ca(OH) ₂
HCl	Sr(OH) ₂
HNO ₃	Ba(OH) ₂

Bronsted, Arrhenius, Lewis

Bronstead	Arrhenius	Lewis
A Bronsted-Lowry acid is anything that releases H ⁺ ions	An Arrhenius acid is a substance that when added to water increases the concentration of H ⁺ ions present	A Lewis acid is therefore any substance, such as the H ⁺ ion, that can accept a pair of nonbonding electrons
A Bronsted-Lowry base is anything that accepts H ⁺ ions	An Arrhenius base is a substance that when added to water increases the concentration of OH ⁻ ions present	A Lewis base is any substance, such as the OH ⁻ ion, that can donate a pair of nonbonding electrons



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Page 1 of 1.

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