

## Acids and Bases

Strong Acids	Strong Bases
HClO <sub>4</sub>	LiOH
H <sub>2</sub> SO <sub>4</sub>	NaOH
HI	KOH
HBr	Ca(OH) <sub>2</sub>
HCl	Sr(OH) <sub>2</sub>
HNO <sub>3</sub>	Ba(OH) <sub>2</sub>

## Bronsted, Arrhenius, Lewis

Bronstead	Arrhenius	Lewis
A Bronsted-Lowry acid is anything that releases H <sup>+</sup> ions	An Arrhenius acid is a substance that when added to water increases the concentration of H <sup>+</sup> ions present	A Lewis acid is therefore any substance, such as the H <sup>+</sup> ion, that can accept a pair of nonbonding electrons
A Bronsted-Lowry base is anything that accepts H <sup>+</sup> ions	An Arrhenius base is a substance that when added to water increases the concentration of OH <sup>-</sup> ions present	A Lewis base is any substance, such as the OH <sup>-</sup> ion, that can donate a pair of nonbonding electrons



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