

### Polyatomic Ions

Ion	Name	Charge
NH <sub>4</sub>	Ammonium	1+
Hg <sub>2</sub>	Mercury(1)	2+
NO <sub>3</sub>	Nitrate	1-
NO <sub>2</sub>	Nitrite	1-
HSO <sub>4</sub>	Bisulfate	1-
OH	Hydroxide	1-
CN	Cyanide	1-
H <sub>2</sub> PO <sub>4</sub>	Dihydrogen Phosphate	1-
NCS	Thiocyanate	1-
HCO <sub>3</sub>	Bicarbonate	1-
ClO	Hypochlorite	1-
ClO <sub>2</sub>	Chlorite	1-
ClO <sub>3</sub>	Chlorate	1-
ClO <sub>4</sub>	Perchlorate	1-
C <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	Acetate	1-
MnO <sub>4</sub>	Permanganate	1-
SO <sub>3</sub>	Sulfite	2-
SO <sub>4</sub>	Sulfate	2-
HPO <sub>4</sub>	Hydrogen Phosphate	2-
CO <sub>3</sub>	Carbonate	2-
Cr <sub>2</sub> O <sub>7</sub>	Dichromate	2-
CrO <sub>4</sub>	Chromate	2-
O <sub>2</sub>	Peroxide	2-
C <sub>2</sub> O <sub>4</sub>	Oxalate	2-
TeO <sub>2</sub>	Hypotellurite	2-
TeO <sub>3</sub>	Tellurite	2-
TeO <sub>4</sub>	Tellurate	2-
TeO <sub>5</sub>	Pertellurate	2-
PO <sub>4</sub>	Phosphate	3-
PO <sub>3</sub>	Phosphite	3-

### Strong Acids

Name	Formula
Hydrochloric Acid	HCl
Hydrobromic Acid	HBr
Hydroiodic Acid	HI
Chloric Acid	HClO <sub>3</sub>
Perchloric Acid	HClO <sub>4</sub>
Nitric Acid	HNO <sub>3</sub>
Sulfuric Acid	H <sub>2</sub> SO <sub>4</sub>

### Strong Bases

Name	Formula
Lithium Hydroxide	LiOH
Sodium Hydroxide	NaOH
Potassium Hydroxide	KOH
Rubidium Hydroxide	RbOH
Calcium Hydroxides	Ca(OH) <sub>2</sub>
Strontium Hydroxide	Sr(OH) <sub>2</sub>

### Other Acids

Name	Formula
Acetic Acid	CH <sub>3</sub> COOH

### Water Energies

Delta H Sublimation	46.6 kJ/mol
Delta H Fusion	6.02 kJ/mol
Delta H Vaporization	40.6 kJ/mol

### Intermolecular Forces

#### Dipole - Dipole Forces

Neutral Polar molecules attract each other when the positive end of one molecule is near the negative end of another. Smaller molecules have a higher dipole-dipole attractive forces.

#### London Dispersion Forces

Not published yet.  
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### Intermolecular Forces (cont)

Neutral non-polar molecules due to the instantaneous distribution of electrons. Temporary dipole on one atom induces a similar dipole on adjacent atom causing the atoms to be attracted to each other.

#### Hydrogen Bonds

An attraction between the hydrogen atom in a polar bond that is bonded to an electronegative atom and the lone pairs of electrons on another atom. Stronger than dipole-dipole or London Dispersion forces.

### Molarity

Molarity (M) = (moles of solute)/(Liters of Solution)

### solubility rules



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