Cheatography

Chemistry Cheat Sheet

by brandonr1950 via cheatography.com/122004/cs/22527/

Polyatomic Ions

Ion	Name	Charge
NH4	Ammonium	1+
Hg2	Mercury(1)	2+
NO3	Nitrate	1-
NO2	Nitrite	1-
HSO4	Bisulfate	1-
OH	Hydroxide	1-
CN	Cyanide	1-
H2PO4	Dihydrogen Phosphate	1-
NCS	Thiocyanate	1-
HCO3	Bicarbonate	1-
ClO	Hypochlorite	1-
ClO2	Chlorite	1-
ClO3	Chlorate	1-
ClO4	Perchlorate	1-
C2H3O2	Acetate	1-
MnO4	Permanganate	1-
SO3	Sulfite	2-
SO4	Sulfate	2-
HPO4	Hydrogen Phosphate	2-
CO3	Carbonate	2-
Cr2O7	Dichromate	2-
CrO4	Chromate	2-
O2	Peroxide	2-
C2O4	Oxalate	2-
TeO2	Hypotellurite	2-
TeO3	Tellurite	2-
TeO4	Tellurate	2-
TeO5	Pertellurate	2-
PO4	Phosphate	3-
PO3	Phosphite	3-

Strong Acids

Name	Formula
Hydrochloric Acid	HCl
Hydrobromic Acid	HBr
Hydroiodic Acid	HI
Chloric Acid	HClO3
Perchloric Acid	HClO4
Nitric Acid	HNO3
Sulfuric Acid	H2SO4

Strong Bases

Name	Formula
Lithium Hydroxide	LiOH
Sodium Hydroxide	NaOH
Potassium Hydroxide	KOH
Rubidium Hydroxide	RbOH
Calcium Hydroxides	Ca(OH)2
Strontium Hydroxide	Sr(OH)2

Other Acids

Name	Formula
Acetic Acid	CH3COOH

Water Energies

Delta H Sublimation	46.6 kJ/mol
Delta H Fusion	6.02 kJ/mol
Delta H Vaporization	40.6 kJ/mol

Intermolecular Forces

Dipole - Dipole Forces

Neutral Polar molecules attract each other when the positive end of one molecule it near the negative end of another. Smaller molecules have a higher dipole-dipole attractive forces.

London Dispersion Forces

Intermolecular Forces (cont)

Neutral non-polar molecules due to the instantaneous distribution of electrons. Temporary dipole on one atom induces a similar dipole on adjacent atom causing the atoms to be attracted to each other.

Hydrogen Bonds

An attraction between the hydrogen atom in a polar bond that is bonded to an electronegative atom and the lone pairs of electrons on another atom. Stronger than dipole-dipole or London Dispersion forces.

Molarity

Molarity (M) = (moles of solute)/(Liters of Solution)

solubility rules

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