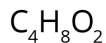


## structural representation

### MOLECULAR FORMULA

The molecular formula of an organic compound simply shows the number of each type of atom present. It tells you nothing about the bonding within the compound.



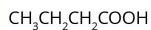
### EMPIRICAL FORMULA

The empirical formula of an organic compound gives the simplest possible whole number ratio of the different types of atom within the compound.



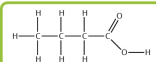
### CONDENSED FORMULA

The condensed formula is also text-based: here, each carbon atom is listed separately, with atoms attached to it following. An exception is cyclic parts of molecules, e.g. benzene, where the carbons are grouped.



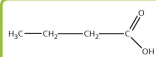
### DISPLAYED FORMULA

A displayed formula shows all of the atoms and all of the bonds present in an organic compound. The bonds are represented as lines.



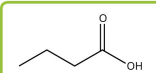
### STRUCTURAL FORMULA

Similar to displayed formula - not all bonds are shown, although all atoms are still indicated using subscript numbers. Carbon hydrogen bonds are often simplified.



### SKELETAL FORMULA

In a skeletal formula, most hydrogen atoms are omitted, and line ends or vertices represent carbons. Functional groups and atoms other than carbon or hydrogen are still shown. Easiest to draw & commonly used.



By beabo  
[cheatography.com/beabo/](https://cheatography.com/beabo/)

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