

### Reaction Rate

- Change in reactants/products' concentration over time

### Factors That Increase Reaction Rate

- Surface area
- Concentration
- Pressure
- Temperature
- Catalysts

### Increasing Above Factors Also Increases...

- Collision frequency
- Successful collisions amount

### Temperature Increase Also Increases...

- Collision energy

### Experiment Methods To Measure Reaction Rate

- Mass loss
- Gas volume made
- Colour change
- Concentration changes
- pH change

### Catalysts

- Provides an alternative reaction pathway with lower activation energy
- Speeds up reaction & isn't used up
- Causes more particle collisions

#### Heterogenous Catalysts

- Has different state to reactants & products

#### Homogenous Catalysts

- Has same state as reactants & products

### Transition State

- Unstable state
- Reactants' bonds broken & products' bonds forming
- Triggered by activation energy

### Collision Theory

- Particle collision affects chemical reaction rates

### Collision Requirements Of Reactant Particles

- Collision with each other
- Collision with enough energy
- Collision in the right direction

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