Chemistry - Chapter 7: Rates Of Chemical Reactions Cheat Sheet

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Reaction Rate

- Change in reactants/products' concentration over time

Factors That Increase Reaction Rate

- Surface area
- Concentration
- Pressure
- Temperature
- Catalysts

Increasing Above Factors Also Increases...

- Collision frequency
- Successful collisions amount

Temperature Increase Also Increases...

- Collision energy

Experiment Methods To Measure Reaction Rate

- Mass loss
- Gas volume made
- Colour change
- Concentration changes
- pH change

Catalysts

- Provides an alternative reaction pathway with lower activation energy
- Speeds up reaction & isn't used up
- Causes more particle collisions

Heterogenous Catalysts	Homogenous Catalysts
Has different state to reactants & producs	Has same state as reactants & products

Transition State

- Unstable state
- Reactants' bonds broken & products' bonds forming
- Triggered by activation energy

Collision Theory

- Particle collision affects chemical reaction rates

Collision Requirements Of Reactant Particles

- Collision with each other
- Collision with enough energy
- Collision in the right direction

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