

CH 1: MATTER, ENERGY, MEASUREMENT

$w = F \times d$	[1.1]	Work done by a force in the direction of displacement
$E_k = \frac{1}{2}mv^2$	[1.2]	Kinetic energy
$K = ^\circ C + 273.15$	[1.3]	Converting between Celsius ($^\circ C$) and Kelvin (K) temperature scales
$^\circ C = \frac{5}{9}(F - 32)$ or $F = \frac{9}{5}(^\circ C) + 32$	[1.4]	Converting between Celsius ($^\circ C$) and Fahrenheit (F) temperature scales
Density = $\frac{\text{mass}}{\text{volume}}$	[1.5]	Definition of density

CH 2: ATOMS, MOLECULES, IONS

Atomic weight = $\sum \left(\frac{\text{isotope mass}}{\text{all isotopes}} \right) \times (\text{fractional isotope abundance})$	[2.1]	Calculating atomic weight as a fractionally weighted average of isotope masses.
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CH 3: CHEM. REACTIONS & REACTION STOICHIOMETRY

This is the formula to calculate the mass percentage of each element in a compound. The sum of all the percentages of all the elements in a compound should add up to 100%.

Mass percentage = $\frac{\text{mass of element}}{\text{molar mass of compound}} \times 100\%$



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Page 1 of 1.

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