The First Law of Thermodynamics

∆(Energy of the System) + ∆(Energy of the Surroundings) = 0

Internal Energy, U

$$\Delta U = U_{final} - U_{initial}$$

Change in Internal Energy

Absorbs heat, Work to Surroundings	∆E = Q - W
Cyclic Process	∆E = 0
Adiabatic Process	W = -∆E ; Q = 0
Change in Internal Energy	∆U = Q + W

PV-Work System

dW = F x dx

P = F / A

dW = P A dx

dV = A dx

dW = P A (dV/A)

W = P dV

Wig = -P (V2-V1)

Wig = nrt (In V1/V2)

By **aes**
cheatography.com/aes/

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